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Answers: Moles and Stoichiometry
Practice Problems 1) How many moles of
sodium atoms correspond to 1.56×10^{21}
atoms of sodium? 1.56×10^{21} atoms Na
 $\times \frac{1 \text{ mol Na}}{6.02 \times 10^{23} \text{ atoms Na}} = 2.59 \times 10^{-3} \text{ mol Na}$

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236.022 x 10 atoms Na 2) Determine the mass in grams of each of the following:

a. 1.35 mol of Fe $1.35 \text{ mol Fe} \times 55.845 \text{ g Fe} = 75.4 \text{ g Fe}$ 1 mol Fe b. 24.5 mol O

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Balance the following chemical reactions: Hint a. $\text{CO} + \text{O}_2 \rightarrow \text{CO}_2$ b. $3 \text{KNO}_3 \rightarrow 3 \text{KNO}_2 + \text{O}_2$ c. $\text{O}_3 \rightarrow \text{O}_2$ d. $2 \text{NH}_4\text{NO}_3 \rightarrow 2 \text{N}_2\text{O} + \text{H}_2\text{O}$ e. $\text{CH}_3\text{NH}_2 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O} + \text{N}_2$ Hint f. $\text{Cr}(\text{OH})_3 + \text{HClO}_4 \rightarrow \text{Cr}(\text{ClO}_4)_3 + \text{H}_2\text{O}$ Write the balanced chemical equations of each reaction:

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Practice Stoichiometry 1 Answer Key

Practice Problems (Chapter 5):

Stoichiometry CHEM 30A Part I: Using the conversion factors in your tool box g A mol A mol A 1. How many moles CH₃OH are in 14.8 g CH₃OH? 2. What is the mass in grams of 1.5 x 10¹⁶ atoms S? 3.

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How many molecules of CO_2 are in 12.0 g CO_2 ? 24. What is the mass in grams of 1 atom of Au? KEY Tool Box: To ...

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Stoichiometry example problem 1.

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Stoichiometry Problem Water is added to
4.267 g of UF₆. The only products are
3.730 g of a solid containing only
uranium, oxygen and fluorine and 0.970
g of a gas. The gas is 95.0% fluorine and
the remainder is hydrogen. (a) From
these data, determine the empirical
formula of the gas.

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Home Lab, Baking Soda Stoichiometry
STEP 1 Fold a sheet of paper in half
lengthwise.

Chapter 11: Stoichiometry

Gas Stoichiometry Practice For all of
these problems, assume that the
reactions are being performed at a
pressure of 1 Gas stoichiometry practice

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answer key. 0 atm and a temperature of 298 K. 1) Calcium carbonate decomposes at high temperatures to form carbon dioxide and calcium oxide:
 $\text{CaCO}_3(\text{s}) \rightarrow \text{CaO}(\text{s}) + \text{CO}_2(\text{g})$
How many grams of calcium carbonate will I need to form 3.45 liters of CO_2

Gas Stoichiometry Practice Answer

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12.3: Mass-Mole and Mole-Mass
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calculations in combination with mole
ratios to solve several different types of
mass-based stoichiometry problems.
Mass to Moles Problems ... but the 1:2
ratio means that more than one mole of

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HF is required for the reaction.
The answer ...

12.3: Mass-Mole and Mole-Mass Stoichiometry - Chemistry ...

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