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Chapter 5 - Electrons in Atoms. Section 5.1 - Models of the

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Atom. The Rutherford's model of the atom did not explain how an atom can emit light or the chemical properties of an atom. Plum Pudding Model Rutherford's Model. The Bohr Model.

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116 Chapter 5 Electrons in Atoms CHAPTER 5 What You'll Learn You will compare the wave and particle models of light. You will describe how the frequency of light emitted by an atom is a unique characteristic of that atom. You will compare and contrast the Bohr and quantum mechanical models of the atom. You will express the arrangements of electrons

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138 Chapter 5 • Electrons in Atoms Although the speed of all electromagnetic waves in a vacuum is the same, waves can have different wavelengths and frequencies. As you can see from the equation on the previous page, wavelength and frequency are inversely related; in other words, as one quantity increases, the other decreases.

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Chapter 5 Electrons in Atoms43 SECTION 5.1 MODELS OF THE ATOM (pages 127-132) This section summarizes the development of atomic theory. It also explains the significance of quantized energies of electrons as they relate to the quantum mechanical model of the atom. The Development of Atomic Models (pages 127-128) 1.

Chapter 5 Electrons In Atoms Test B Answers

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Chapter 5 - Electrons in Atoms - 5 Assessment - Page 154: 80
Answer Electrons occupy equal energy orbitals singly before beginning to pair up, because according to Hund's rule the most stable arrangement of electrons occurs when electrons with the same spin is as large as possible.

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Glencoe Chemistry: Matter and Change Buthelezi, Dingrando, Wistrom, Zike. Chapter 5 Electrons in Atoms. Educators. AY IB Chapter Questions. 02:25. Problem 1 Objects get their colors from reflecting only certain wavelengths when hit with white light. Light reflected from a green leaf is found to have a wavelength of $4.90 \times 10^{-7} \text{ m}$

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